

will be : **(A)**

0.2 M

(B)

10.3 M

solution acidic because: (A) It does not react with KMnO ₄ or the reducing agent (B) Hydrochloric acid reacts with KMnO ₄ (C) Nitric acid is an oxidising agent which reacts with reducing agent (D) All of the above are correct (D) All of the above are correct (D) All of the above are correct (E) The compound YBa ₂ Cu _n O ₇ has copper in oxidation state +1. Assume that the rare earth eler Yttrium is in its usual +3 oxidation state. The value of n is: (A) 3/7 (B) 7/3 (C) 3 (D) 7 (Or Question No. 88 - 89 (A) Statement-1 is True, Statement-2 is True and Statement-2 is NOT a correct explanation for Statement-1 is True, Statement-2 is True and Statement-2 is NOT a correct explanation for Statement Statement-1 is False, Statement-1 is True. Statement-1 is False, Statement-2 is True. Statement-1: Stannous chloride is a powerful oxidising agent which oxidises mercuric chloric mercury. Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chlored does not do so. Statement 1: H ₂ SO ₄ cannot act as reducing agent. Statement 2: Sulphur cannot increase its oxidation number beyond + 6. Among the following, identify the species with an atom in +6 oxidation state. (A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ (A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ (A) M/2 (B) M/8 (C) M/3 (D) M (D)	Date Planned ://				Daily	Daily Tutorial Sheet-7			Expected Duration : 90 Min		
solution acidic because: (A) It does not react with KMnO ₄ or the reducing agent (B) Hydrochloric acid reacts with KMnO ₄ (C) Nitric acid is an oxidising agent which reacts with reducing agent (D) All of the above are correct (T) The compound YBa ₂ Cu _n O ₇ has copper in oxidation state +1. Assume that the rare earth eler Yttrium is in its usual +3 oxidation state. The value of n is: (A) 3/7 (B) 7/3 (C) 3 (D) 7 For Question No. 88 - 89 (A) Statement-1 is True, Statement-2 is True and Statement-2 is a correct explanation for Statement-1. Statement-1 is True, Statement-2 is True and Statement-2 is NOT a correct explanation for Statement-1. Statement-1 is True, Statement-2 is True. Statement-1 is False, Statement-2 is True. Statement-1 is False, Statement-2 is True. Statement-1: Stannous chloride is a powerful oxidising agent which oxidises mercuric chloric mercury. Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chlodoes not do so. Statement 2: Sulphur cannot increase its oxidation number beyond + 6. Anong the following, identify the species with an atom in +6 oxidation state. (A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ In the estimation of Na ₂ S ₂ O ₃ using Br ₂ the equivalent weight of Na ₂ S ₂ O ₃ is: (A) M/2 (B) M/8 (C) M/3 (D) M 13. A 0.1097 gm sample of As ₂ O ₃ required 26.10 mL of KMnO ₄ solution for its titration. The molari KMnO ₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 44. In basic medium, CrO ₄ ² oxidizes S ₂ O ₃ ² to form SO ₄ ² and itself changes to Cr(OH) ₄ . How many no.154 M CrO ₄ ² are required to react with 40 mL of 0.246 M S ₂ O ₃ ² ? (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL	Actu	al Date	of Attempt : _	_/_/_		Level-2			Exact Duration :		
(D) All of the above are correct 37. The compound YBa ₂ Cu _n O ₇ has copper in oxidation state +1. Assume that the rare earth elect Yttrium is in its usual +3 oxidation state. The value of n is: (A) 3/7 (B) 7/3 (C) 3 (D) 7 Statement 1 is True, Statement 2 is True and Statement 2 is a correct explanation for Statement 1. Statement 1 is True, Statement 2 is True and Statement 2 is NOT a correct explanation for Statement Statement 1 is True, Statement 2 is True. Statement 1 is True, Statement 2 is True. Statement 1: Stannous chloride is a powerful oxidising agent which oxidises mercuric chloric mercury. Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chlorides not do so. Statement 1: H ₂ SO ₄ cannot act as reducing agent. Statement 2: Sulphur cannot increase its oxidation number beyond + 6. Among the following, identify the species with an atom in +6 oxidation state. (A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ 31. In the estimation of Na ₂ S ₂ O ₃ using Br ₂ the equivalent weight of Na ₂ S ₂ O ₃ is: (A) M/2 (B) M/8 (C) M/3 (D) M 32. In the reaction 3ClO (aq) → ClO ₃ (aq) + 2Cl (aq) the equivalent mass of ClO is: (A) M/4 (B) M/3 (C) 3M (D) 3M/4 33. A 0.1097 gm sample of As ₂ O ₃ required 26.10 mL of KMnO ₄ solution for its titration. The molari KMnO ₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 34. In basic medium, CrO ₄ ² oxidizes S ₂ O ₃ ² to form SO ₄ ² and itself changes to Cr(OH) ₄ . How many nounce of the content of the co	36.	solutio	on acidic becaus	se : act with KM	${ m inO_4}$ or the re			und to be	e suitable for r	naking the	
Yttrium is in its usual +3 oxidation state. The value of n is: (A) 3/7 (B) 7/3 (C) 3 (D) 7 For Question No. 88 - 89 A) Statement-1 is True, Statement-2 is True and Statement-2 is a correct explanation for Statement-1. Statement-1 is True, Statement-2 is True and Statement-2 is NOT a correct explanation for Statement Statement-1 is True, Statement-2 is False. Statement-1 is True, Statement-2 is True. Statement-1 is False, Statement-2 is True. Statement-1 is False, Statement-2 is True. Statement 1: Stannous chloride is a powerful oxidising agent which oxidises mercuric chloride mercury. Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chloride does not do so. Statement 2: Sulphur cannot increase its oxidation number beyond + 6. Among the following, identify the species with an atom in +6 oxidation state. (A) Ba₂XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO₂Cl₂ In the estimation of Na₂S₂O₃ using Br₂ the equivalent weight of Na₂S₂O₃ is: (A) M/2 (B) M/8 (C) M/3 (D) M 12. In the reaction 3ClO⁻(aq) → ClO₃(aq) + 2Cl⁻(aq) the equivalent mass of ClO⁻ is: (A) M/4 (B) M/3 (C) 3M (D) 3M/4 33. A 0.1097 gm sample of As₂O₃ required 26.10 mL of KMnO₄ solution for its titration. The molari KMnO₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 44. In basic medium, CrO₄² oxidizes S₂O₃² to form SO₄² and itself changes to Cr(OH)₄. How many nolification is: (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL				_	_	eacts w	th reducing a	gent			
Statement-1 is True, Statement-2 is True and Statement-2 is a correct explanation for Statement-1. Statement-1 is True, Statement-2 is True and Statement-2 is NOT a correct explanation for Statement Statement-1 is True, Statement-2 is False. Statement-1 is True, Statement-2 is True. Statement-1 is False, Statement-2 is True. Statement 1: Stannous chloride is a powerful oxidising agent which oxidises mercuric chloride mercury. Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chloroses not do so. Statement 1: H ₂ SO ₄ cannot act as reducing agent. Statement 2: Sulphur cannot increase its oxidation number beyond + 6. Among the following, identify the species with an atom in +6 oxidation state. (A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ In the estimation of Na ₂ S ₂ O ₃ using Br ₂ the equivalent weight of Na ₂ S ₂ O ₃ is: (A) M/2 (B) M/8 (C) M/3 (D) M In the reaction 3ClO ⁻ (aq) — ClO ⁻ ₃ (aq) + 2Cl ⁻ (aq) the equivalent mass of ClO ⁻ is: (A) M/4 (B) M/3 (C) 3M (D) 3M/4 A 0.1097 gm sample of As ₂ O ₃ required 26.10 mL of KMnO ₄ solution for its titration. The molari KMnO ₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 A In basic medium, CrO ² ₄ oxidizes S ₂ O ² ₃ to form SO ² ₄ and itself changes to Cr(OH) ₄ . How many nounce the content of th	37.	Yttriu	m is in its usua	l +3 oxidati	ion state. The	value of	n is:			th elemen	
Statement-1 is True, Statement-2 is True and Statement-2 is a correct explanation for Statement-1. Statement-1 is True, Statement-2 is True and Statement-2 is NOT a correct explanation for Statement Statement-1 is True, Statement-2 is False. Statement-1 is True, Statement-2 is True. Statement-1 is False, Statement-2 is True. Statement-1 is False, Statement-2 is True. Statement 1: Stannous chloride is a powerful oxidising agent which oxidises mercuric chloride mercury. Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chlorides not do so. Statement 1: H ₂ SO ₄ cannot act as reducing agent. Statement 2: Sulphur cannot increase its oxidation number beyond + 6. Among the following, identify the species with an atom in +6 oxidation state. (A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ In the estimation of Na ₂ S ₂ O ₃ using Br ₂ the equivalent weight of Na ₂ S ₂ O ₃ is: (A) M/2 (B) M/8 (C) M/3 (D) M In the reaction 3ClO ⁻ (aq) — ClO ⁻ ₃ (aq) + 2Cl ⁻ (aq) the equivalent mass of ClO ⁻ is: (A) M/4 (B) M/3 (C) 3M (D) 3M/4 A 0.1097 gm sample of As ₂ O ₃ required 26.10 mL of KMnO ₄ solution for its titration. The molari KMnO ₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 In basic medium, CrO ² ₄ oxidizes S ₂ O ² ₃ to form SO ² ₄ and itself changes to Cr(OH) ₄ . How many nounce the content of th	or Q	uestion	No. 88 - 89								
Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chlorides not do so. 39. Statement 1: H_2SO_4 cannot act as reducing agent. Statement 2: Sulphur cannot increase its oxidation number beyond + 6. 30. Among the following, identify the species with an atom in +6 oxidation state. (A) Ba_2XeO_6 (B) XeO_3 (C) CrO_5 (D) SO_2Cl_2 31. In the estimation of $Na_2S_2O_3$ using Br_2 the equivalent weight of $Na_2S_2O_3$ is: (A) $M/2$ (B) $M/8$ (C) $M/3$ (D) M 32. In the reaction $3ClO^-(aq) \longrightarrow ClO_3^-(aq) + 2Cl^-(aq)$ the equivalent mass of ClO^- is: (A) $M/4$ (B) $M/3$ (C) $3M$ (D) $3M/4$ 33. A 0.1097 gm sample of As_2O_3 required 26.10 mL of $KMnO_4$ solution for its titration. The molari $KMnO_4$ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 34. In basic medium, CrO_4^2 oxidizes $S_2O_3^2$ to form SO_4^2 and itself changes to $Cr(OH)_4$. How many $0.154 M CrO_4^2$ are required to react with 40 mL of $0.246 M S_2O_3^2$? (A) $0.00 ML$ (B) $156.4 ML$ (C) $170.4 ML$ (D) $190.4 ML$	B) C)	Statement-1 is True, Statement-2 is True and Statement-2 is NOT a correct explanation for Statement-1 Statement-1 is True, Statement-2 is False.									
Statement 2: Sulphur cannot increase its oxidation number beyond + 6. 90. Among the following, identify the species with an atom in +6 oxidation state. (A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ 91. In the estimation of Na ₂ S ₂ O ₃ using Br ₂ the equivalent weight of Na ₂ S ₂ O ₃ is: (A) M/2 (B) M/8 (C) M/3 (D) M 12. In the reaction 3 ClO ⁻ (aq) \longrightarrow ClO ₃ ⁻ (aq) + 2Cl ⁻ (aq) the equivalent mass of ClO ⁻ is: (A) M/4 (B) M/3 (C) 3M (D) 3M/4 13. A 0.1097 gm sample of As ₂ O ₃ required 26.10 mL of KMnO ₄ solution for its titration. The molarity KMnO ₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 14. In basic medium, C rO ₄ ⁻ oxidizes S ₂ O ₃ ⁻ to form SO ₄ ⁻ and itself changes to C r(OH) ₄ . How many molarity MCrO ₄ ⁻ are required to react with 40 mL of 0.246 M S ₂ O ₃ ⁻ ? (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL	88.	Statement 2: Stannous chloride gives grey precipitate with mercuric chloride, but stannic chloride									
Among the following, identify the species with an atom in $+6$ oxidation state. (A) Ba_2XeO_6 (B) XeO_3 (C) CrO_5 (D) SO_2Cl_2 1. In the estimation of $Na_2S_2O_3$ using Br_2 the equivalent weight of $Na_2S_2O_3$ is: (A) $M/2$ (B) $M/8$ (C) $M/3$ (D) M 12. In the reaction $3ClO^-(aq) \longrightarrow ClO_3^-(aq) + 2Cl^-(aq)$ the equivalent mass of ClO^- is: (A) $M/4$ (B) $M/3$ (C) $3M$ (D) $3M/4$ 13. A 0.1097 gm sample of As_2O_3 required 26.10 mL of $KMnO_4$ solution for its titration. The molari $KMnO_4$ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 14. In basic medium, CrO_4^{2-} oxidizes $S_2O_3^{2-}$ to form SO_4^{2-} and itself changes to $Cr(OH)_4^-$. How many 0.154 M CrO_4^{2-} are required to react with 40 mL of 0.246 M $S_2O_3^{2-}$? (A) 0.00 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL	9.	State	nent 1: H ₂ SO	O ₄ cannot a	ct as reducing	agent.					
(A) Ba ₂ XeO ₆ (B) XeO ₃ (C) CrO ₅ (D) SO ₂ Cl ₂ 91. In the estimation of Na ₂ S ₂ O ₃ using Br ₂ the equivalent weight of Na ₂ S ₂ O ₃ is: (A) M/2 (B) M/8 (C) M/3 (D) M 92. In the reaction 3ClO ⁻ (aq) → ClO ₃ ⁻ (aq) + 2Cl ⁻ (aq) the equivalent mass of ClO ⁻ is: (A) M/4 (B) M/3 (C) 3M (D) 3M/4 93. A 0.1097 gm sample of As ₂ O ₃ required 26.10 mL of KMnO ₄ solution for its titration. The molari KMnO ₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 94. In basic medium, CrO ₄ ⁻ oxidizes S ₂ O ₃ ²⁻ to form SO ₄ ⁻ and itself changes to Cr(OH) ₄ . How many nounce the content of		State	nent 2: Sulpl	nur cannot ii	ncrease its ox	dation r	number beyon	1 + 6.			
 (A) M/2 (B) M/8 (C) M/3 (D) M (D) M (E) M/4 (B) M/3 (C) 3M (D) 3M/4 (E) M/4 (D) 3M/4 (E) M/4 (D) 3M/4 (E) M/4 (D) 3M/4 (E) M/5 (D) 3M/4 (E) M/6 (D) M/6 (E) M/6 (D) M/7 (E) M/7 (E) M/8 (D) M/8 (E) M	90.	•			_				SO_2Cl_2		
(A) M/4 (B) M/3 (C) 3M (D) 3M/4 33. A 0.1097 gm sample of As_2O_3 required 26.10 mL of $KMnO_4$ solution for its titration. The molari $KMnO_4$ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 34. In basic medium, CrO_4^{2-} oxidizes $S_2O_3^{2-}$ to form SO_4^{2-} and itself changes to $Cr(OH)_4^{-}$. How many nounce 0.154 M CrO_4^{2-} are required to react with 40 mL of 0.246 M $S_2O_3^{2-}$? (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL	91.						-		M		
A 0.1097 gm sample of As_2O_3 required 26.10 mL of $KMnO_4$ solution for its titration. The molari $KMnO_4$ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 1. In basic medium, CrO_4^{2-} oxidizes $S_2O_3^{2-}$ to form SO_4^{2-} and itself changes to $Cr(OH)_4^{-}$. How many nounce 0.154 M CrO_4^{2-} are required to react with 40 mL of 0.246 M $S_2O_3^{2-}$? (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL	92.	In the	reaction 3ClO	(aq) (ClO ₃ (aq) + 2 Cl	(aq) th	e equivalent m	ass of Clo	O is:		
KMnO ₄ solution is: (A) 0.02 (B) 0.04 (C) 0.018 (D) 0.3 1. In basic medium, CrO_4^{2-} oxidizes $S_2O_3^{2-}$ to form SO_4^{2-} and itself changes to $Cr(OH)_4^-$. How many none of 0.154 M CrO_4^{2-} are required to react with 40 mL of 0.246 M $S_2O_3^{2-}$? (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL		(A)	M/4	(B)	M/3	(C)	3 M	(D)	3M/4		
10.154 M $\text{CrO}_4^{2^-}$ are required to react with 40 mL of 0.246 M $\text{S}_2\text{O}_3^{2^-}$? (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL	93.	KMnC	0_4 solution is :							molarity o	
0.154 M CrO_4^{2-} are required to react with 40 mL of 0.246 M $S_2O_3^{2-}$? (A) 200 mL (B) 156.4 mL (C) 170.4 mL (D) 190.4 mL											
	4.	0.154	$\mathrm{M}\;\mathrm{CrO}_4^{2-}\;\mathrm{are}\;\mathrm{re}$	equired to re	eact with 40 m	L of 0.2	46 M S ₂ O ₃ ²⁻ ?			nany mL o	
95. 20 mL of $0.2 \mathrm{MAl_2(SO_4)_3}$ is mixed with 20 mL of $0.6 \mathrm{MBaCl_2}$. Concentration of $\mathrm{Al^{3+}}$ ion in the solu											

VMC | Stoichiometry-II 160 DTS-7 | Level-2

(C)

 $0.1~\mathrm{M}$

(D)

 $0.25 \; M$

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